

M. Khawar Rauf,^{a*} Amin
Badshah^a and Michael Bolte^b

^aDepartment of Chemistry, Quaid-i-Azam
University, Islamabad 45320, Pakistan, and

^bInstitut für Anorganische Chemie, J. W.
Goethe-Universität Frankfurt, Max-von-Laue-
Strasse 7, 60438 Frankfurt/Main, Germany

Correspondence e-mail:
khawar_rauf@hotmail.com

Key indicators

Single-crystal X-ray study
 $T = 173$ K
Mean $\sigma(\text{C}-\text{C}) = 0.003$ Å
 R factor = 0.034
 wR factor = 0.087
Data-to-parameter ratio = 14.3

For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.

1-(3-Chlorobenzoyl)-3-(2,4,6-trichlorophenyl)- thiourea

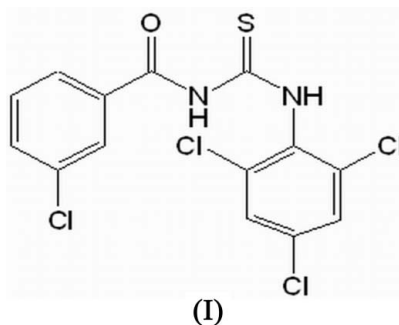
The title compound, $\text{C}_{14}\text{H}_8\text{Cl}_4\text{N}_2\text{OS}$, shows the typical geometric parameters of substituted thiourea derivatives. Whereas one NH group is involved in an intramolecular hydrogen bond, the other is shielded, so that no hydrogen bond is observed.

Received 15 May 2006

Accepted 18 May 2006

Comment

N -substituted and N,N' -disubstituted thiourea derivatives have potential applications due to their coordination behaviour towards transition metals (Schuster *et al.*, 1990) and their biological activity (Frech *et al.*, 1970; Madan *et al.*, 1991).



The title compound, (I), (Fig. 1) is a typical N,N' -disubstituted thiourea derivative with normal geometric parameters (Cambridge Structural Database, version 5.27 plus one update; *MOGUL* Version 1.7; Allen, 2002). The C2–S1 and C1–O1 bonds (Table 1) both show the expected full double-bond character, while the short values for the C1–N1, C2–N1, C2–N2 and C21–N2 bond lengths indicate partial double-bond character. The dihedral angle between the aromatic rings is $84.17(6)^\circ$. The dihedral angle between the thiourea plane (O1/C1/N1/C2/S1/N2) and the ring formed by atoms C11–C16 is $3.86(13)^\circ$, and that between the thiourea plane and the ring formed by atoms C21–C26 is $86.58(5)^\circ$.

An intramolecular N–H \cdots O hydrogen bond is present (Table 2), which results in the formation of a six-membered ring, as is commonly observed in most benzoylthiourea derivatives (Arslan *et al.*, 2004; Khawar Rauf *et al.*, 2006). The other NH group is shielded, so that no hydrogen-bond acceptor can approach it to form an intermolecular hydrogen bond. A weak hydrogen bond between C23–H of an aromatic ring and S is noted (Table 2).

Experimental

Freshly prepared 3-chlorobenzoyl chloride (1.75 g, 10 mmol) was added to a suspension of KSCN ((1.00 g, 10 mmol) in acetone

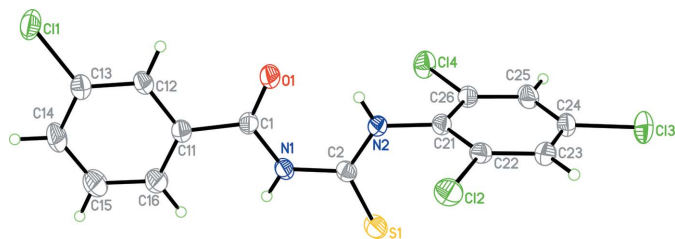


Figure 1

The molecular structure of (I). Displacement ellipsoids are drawn at the 50% probability level.

(30 ml). The reaction mixture was stirred for 15 min. Neat 2,4,6-trichloroaniline (1.97 g, 10 mmol) was then added and the resulting mixture was stirred for 1 h. The reaction mixture was then poured into acidified water and stirred well. The solid product was separated, washed with deionized water and purified by recrystallization from methanol–dichloromethane (1:1), to give fine crystals of (I) (overall yield 85%; m.p. 453 K).

Crystal data

$C_{14}H_8Cl_4N_2OS$
 $M_r = 394.08$
 Monoclinic, $P2_1/n$
 $a = 12.4021$ (12) Å
 $b = 9.2723$ (9) Å
 $c = 14.5212$ (15) Å
 $\beta = 107.379$ (8)°
 $V = 1593.6$ (3) Å³

$Z = 4$
 $D_x = 1.643$ Mg m⁻³
 Mo $K\alpha$ radiation
 $\mu = 0.87$ mm⁻¹
 $T = 173$ (2) K
 Block, colourless
 $0.48 \times 0.46 \times 0.43$ mm

Data collection

Stoe IPDS-II two-circle
 diffractometer
 ω scans
 Absorption correction: multi-scan
 (MULABS; Spek, 2003; Blessing,
 1995)
 $T_{\min} = 0.679$, $T_{\max} = 0.705$

8381 measured reflections
 2968 independent reflections
 2559 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.046$
 $\theta_{\text{max}} = 25.6^\circ$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.087$
 $S = 1.03$
 2968 reflections
 208 parameters
 H atoms treated by a mixture of
 independent and constrained
 refinement

$w = 1/[\sigma^2(F_o^2) + (0.0446P)^2 + 0.8237P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} = 0.001$
 $\Delta\rho_{\text{max}} = 0.29$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.28$ e Å⁻³
 Extinction correction: SHELXL97
 (Sheldrick, 1997)
 Extinction coefficient: 0.0115 (11)

Table 1

Selected bond lengths (Å).

S1—C2	1.662 (2)	C1—O1	1.226 (3)
C11—C13	1.750 (2)	C1—N1	1.372 (3)
C12—C22	1.737 (2)	C2—N2	1.340 (3)
C13—C24	1.743 (2)	N2—C21	1.426 (3)
C14—C26	1.738 (2)		

Table 2

Hydrogen-bond geometry (Å, °).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
N2—H2 \cdots O1	0.82 (3)	2.01 (3)	2.650 (2)	135 (3)
C23—H23 \cdots S1 ⁱ	0.95	2.74	3.638 (2)	158

Symmetry code: (i) $-x + \frac{1}{2}, y - \frac{1}{2}, -z + \frac{1}{2}$.

C-bound H atoms were included in the riding-model approximation, with C—H = 0.95 Å and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$. N-bound H atoms were refined freely; N—H = 0.82 (3) Å.

Data collection: X-AREA (Stoe & Cie, 2001); cell refinement: X-AREA; data reduction: X-AREA; program(s) used to solve structure: SHELXS97 (Sheldrick, 1990); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: XP in SHELXTL-Plus (Sheldrick, 1991); software used to prepare material for publication: SHELXL97.

MKR is grateful to the Higher Education Commission of Pakistan for financial support for a PhD programme [scholarship No. (PIN) ILC (0363104)].

References

- Allen, F. H. (2002). *Acta Cryst.* **B58**, 380–388.
 Arslan, H., Flörke, U. & Külcü, N. (2004). *Turk. J. Chem.* **28**, 673–678.
 Blessing, R. H. (1995). *Acta Cryst.* **A51**, 33–38.
 Frech, F. A., Blanz, E. J., Amaral, J. R. D. & French, D. A. (1970). *J. Med. Chem.* **13**, 1117–1124.
 Khawar Rauf, M., Badshah, A., Saeed, A. & Bolte, M. (2006). *Acta Cryst.* **E62**, o1262–o1263.
 Madan, V. K., Taneja, A. D. & Kudesia, V. P. (1991). *J. Indian Chem. Soc.* **68**, 471–472.
 Schuster, M., Kugler, B. & König, K. H. (1990). *J. Anal. Chem.* **338**, 717–720.
 Sheldrick, G. M. (1990). *Acta Cryst.* **A46**, 467–473.
 Sheldrick, G. M. (1991). *SHELXTL-Plus*. Release 4.1. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
 Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.
 Spek, A. L. (2003). *J. Appl. Cryst.* **A36**, 7–13.
 Stoe & Cie (2001). *X-AREA*. Stoe & Cie, Darmstadt, Germany.